Acknowledgment

The authors wish to acknowledge their indebtedness to the late Professor Victor Lenher of the University of Wisconsin and to Dr. S. J. Dickinson of the American Smelting and Refining Company for valuable suggestions as to the precipitation of selenium from selenates by the potassium iodide method.

Summary

The Parr bomb fusion method has been adapted to the analysis of organic selenium compounds. The conditions for the precipitation of the selenium from the solution of the melt have been worked out.

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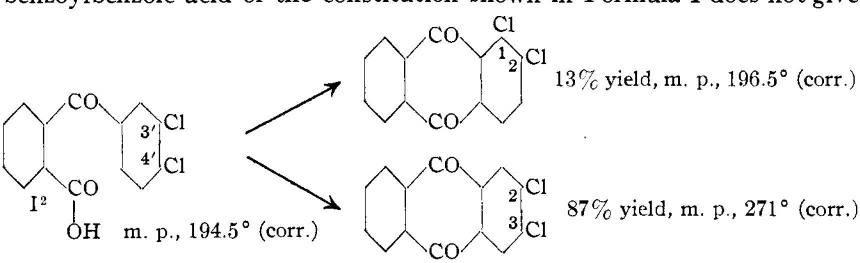
NOTES

The Formation of 1,2-Dichloro-anthraquinone and 2,3-Dichloro-anthraquinone from *o*-Dichlorobenzene.—Phillips states[1] that dichlorobenzoylbenzoic acid of the constitution shown in Formula I does not give



any 1,2-dichloro-anthraquinone on condensation with concentrated sulfuric acid. This statement is erroneous as I have shown[3] that there is always formed about 13% of 1,2-dichloro-anthraquinone, which remains in the sulfuric acid after the condensation and crystallizes on diluting the acid in little leaflets. The substance melts, contrary to the statement of Ullmann,[4] at 196.5° (corr.), and not at 207°.

I may add that Phillips' observations are, on the other hand, quite in agreement with ours. The 1,2-dichlorobenzene was quite pure, being especially made for us by the Aktiengesellschft für Aniline Fabrikation in Berlin. I have already published this fact on page 488 of my "Künstliche Organische Farbstoffe" (Berlin, 1926). The yield of the 3′,4′-dichlorobenzoylbenzoic acid was in our best experiments 35%.[5]

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RECEIVED JUNE 2, 1927
PUBLISHED SEPTEMBER 2, 1927

[1] Phillips, THIS JOURNAL, **49**, 473–478 (1927).

[2] The sodium salt as well as the potassium salt is sparingly soluble.

[3] E. Senn, *Dissertation*, Swiss Technical High School, **1923**.

[4] See Ullmann, Encyklopaedia.

[5] Ref. 3, p. 29.